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# **Chapter 12**

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can



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determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

### **Chapter 12.2: Stoichiometry of Reactions in Solution ...**

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Practice Problems Answer Key A In any stoichiometry problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are Page 4/21. File Type PDF Chapter 12 Stoichiometry Practice Problems Answers given the mass of K<sub>2</sub>

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Analysis Data Analysis 1. A sample of  
aluminum is placed in a 25-mL  
graduated cylinder containing 10.0 mL

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of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample. 2. Saturn is about 1 429 000 km from the Sun.

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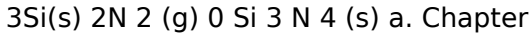
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